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A SCHEME FOR SEPARATING AND DETECTING SULPHIDES, CYANIDES, FERROCYANIDES, FERRICYANIDES, CHLORIDES, BROMIDES, AND IODIDES IN THE PRESENCE OF EACH OTHER.

By J. S. C. Wells and H. T. Vulte.*

Boil out H_2S, test fumes with PbA_2 paper a black coloration indicates H_2S. Add $NaHCO_3$ and distil, carrying distillation to dryness, and pass distillate into KOH, test this for HCy in usual way, when cyanides are removed, cool, dissolve residue in H_2O and add $CuSO_4$ as long as a ppt. is formed, and until the liquid becomes blue or green ; filter.

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* *School of Mines Quarterly*, *X.*, 3.

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Ppt.	Soln.
Cu_2 Fe Cy_6 Cu_3 Fe_2 Cy_{12} NOTE— Test for these acids in the original with Ferric and Ferrous salts.	Add conc. solution of $(NH_4)_2$ Fe_2 $(SO_4)_4$ and boil, pass vapour into starch water. Blue colour indicates HI when this is removed add a drop of conc. H_2SO_4 some dil. $K_2Mn_2O_8$ and boil as long as Br is given off or until the colour is discharged, adding more $K_2Mn_2O_8$ if necessary. Finally test soln. with HNO_3 and $AgNO_3$, a white curdy ppt. sol. in NH_4OH and pptd. by HNO_3 indicates HCl.

NOTE.—So much of the above scheme as relates to separation of Cl Br and I is due to E. Hart (See *Chem. News*, Dec., 1884).

The use of $(NH_4)_2$ Fe_2 $(SO_4)_4$ is original.